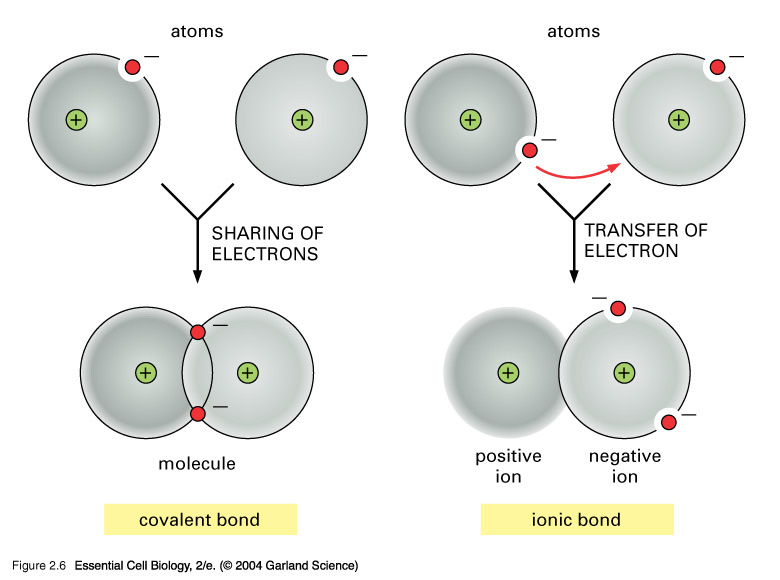
**Ionic Covalent Bond**

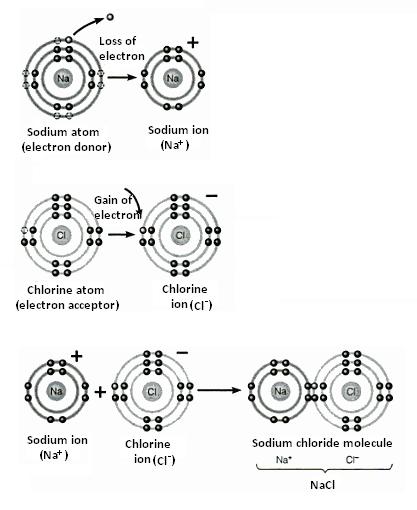
**Vocabulary**

* **Anion:** An ion that has a negative charge.
* **Atom:** The smallest unit of an element that retains the chemical properties of the element. Atoms can exist alone or in combinations with other atoms forming molecules.
* **Cation:** An ion that has a positive charge.
* **Chemical Bond:** A link between atoms.
* **Compound:** A substance formed by the chemical combination of elements in defined proportions.
* **Covalent Bond:** A chemical bond formed by the sharing of electrons between atoms.
* **Dipole:** The uneven distribution of electrical charge across a polar molecule.
* **Electronegativity:** A measure of the attraction of an atom for the electrons in a covalent bond.
* **Ion:** An atom that has acquired an electrical charge by either gaining or losing electrons.
* **Ionic Bond:** Bond formed between oppositely charged ions because of electrostatic forces of attraction.
* **Ionic Compound:** A compound formed by ions.
* **Lewis Dot Structure:** A diagram to represent the valence electrons of an atom, written as the element symbol surrounded by dots that represent the valence electrons. Also used to represent bonding between atoms.
* **Molecule:** A unit of matter formed by the chemical bonding of two or more atoms by covalent bonds.
* **Neutral Atom:** An atom with equal numbers of protons and electrons.
* **Octet Rule (Noble Gas Rule):** Rule that states that atoms tend to combine in such a way that they each have eight electrons in the highest main energy level, giving them the same electronic configuration as a noble gas.
* **Valence Electron:** Electrons in the outermost energy level of an atom; Electrons that can be involved in chemical bonding.

**Ionic Covalent Bond**

**Diagram Configuration**





**Ionic Bond Salt NaCl**

